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Unit 3 Atomic Structure- Guided Notes
The Periodic Table ... - Add you answers from number 1 together - Round your answer to 2 decimal places • Practice: 1. A sample of Cesium is 75% ^{133}Cs , 20% ^{132}Cs , and 5% ^{134}Cs . What is the average atomic mass? 2. Carbon-12 is 98% abundant and Carbon-14 is 2% abundant.

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Structure- Guided Notes The Periodic Table _____ is credited for creating the first version of the modern periodic table _____ is a chart showing all the elements arranged in columns in such a way that all the elements in a given column exhibit similar chemical properties The periodic table is organized in order of _____ The periodic is organized in&mldr ...

Unit 2 - Atomic Structure Guided Notes 2020.docx - Name ...

neutrons in any given element, the atomic number (or number of protons) is subtracted from the mass number. An example: A lithium atom has the atomic number (Z) of 3, and the atomic mass (A) of 6.94 (round it up to 7!). Therefore lithium contains three protons (found from Z), and

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Guided Notes: Atomic Structure 01 Gold Foil Experiment Sketch and label the

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Rutherford Gold Foil Experiment. In this model, where are each of the subatomic particles? o Proton - o Neutron - Distinguishing among Elements Elements differ in their number of _____.

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- Write the atomic mass at the bottom of the square. Step 2: Determine the number of protons, neutrons, and electrons in each element. Step 3: Create an atomic structure diagram for each element. Step 4: Draw the Bonding Structure for each element. 5 B Boron 10.81 P = 5 N = 6 E = 5 Atomic Structure: Bonding Structure: B 5 P+ 6 no

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MASS NUMBER: The sum of the protons and neutrons in the nucleus of an atom.
 $MASS\# = \# \text{ protons(atomic number)} + \# \text{neutrons.} \#NEUTRONS = Mass\# - \#Protons(atomic\ number) \#Protons = Mass\# - \#Neutrons.$ ISOTOPEs: Are atoms of the same element that have different numbers of. neutrons and different mass numbers.

Chapter 4: Atomic Structure Notes Flashcards | Quizlet

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Worksheet: Chapter 4 - Atomic Structure, Sheet 1 HW: Worksheet: Chapter 4 - Atomic Model Timeline. Day 2 - IPOD # 7 - atomic structure chart Review HW Worksheet: Atomic Structure, Sheet 1 (finish) Chapter 4 Notes, Slides 14-16: Effect of changing particles Chapter 4 Notes, Slide 17: Isotope definition Chapter 4 Notes, Slide 18-20: Practice

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Unit 3: Atoms & Elements - Chemistry by Mrs. Hinkson

Atoms Guided Answers 6 88 mass number 4 5 neon-22 10 12 10 22 Ne 10 mass ... 5 ATOMIC STRUCTURE AND THE PERIODIC TABLE The atom is a ball of positive charge with electrons stuck into the ball. Electrons are arranged in concentric circular paths around the nucleus. Most of an atom's mass is

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concentrated in the small, positively charged nucleus. The electrons

Electrons In Atoms Guided Answers

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